

PO LEUNG KUK CENTENARY LI SHIU CHUNG MEMORIAL COLLEGE
TEACHING SCHEDULE
FORM 3 CHEMISTRY
(2023-2024)

CYCLE	DATES	SYLLABUS TO BE COVERED	REMARKS
1	4/9 - 12/9	Unit 1 Introducing Chemistry 1.1 What is Chemistry? 1.2 Why study Chemistry? 1.3 Laboratory safety 1.4 Hazard warning labels 1.5 Common apparatus in the laboratory	Ex.1
2	13/9 - 20/9	Unit 2 The atmosphere 2.1 The Earth 2.2 Classification of matter 2.3 Elements and compounds 2.4 Differences between a mixture and a compound 2.6 The atmosphere 2.7 Physical and Chemical Properties 2.8 Physical and chemical changes 2.9 Separation of mixtures 2.10 Separating oxygen and nitrogen from the air 2.11 Test for oxygen	Ex.2
3	21/9 - 28/9	Unit 3 The ocean 3.1 Sea water: a vast solution 3.2 What is a solution? 3.3 Obtaining common salt from sea water 3.4 Obtaining pure water from sea water Experiment 3.1	
4	29/9 - 9/10	Unit 3 The ocean 3.5 What does common salt contain? 3.6 Test for presence of water in a sample 3.7 Composition of sea water 3.8 Getting useful substances from sea water Experiment 3.2	Ex.3
5	10/10 - 17/10	Unit 4 : Rocks and minerals 4.1 Metals in the Earth's crust 4.2 Extracting metals from their ores 4.3 Investigating calcium carbonate	
6	18/10 - 1/11	Unit 4 The ocean 4.4 Formation of chalk, limestone and marbles 4.5 Formation of limestone caves Experiment 4.1	Ex.4 1st Uniform test
7	2/11 - 13/11	Unit 5 Atomic structure 5.1 What is an element made of? 5.2 Symbols for elements 5.3 States of elements 5.4 How to classify elements? 5.5 Basic Structure of an atom 5.6 Atomic number 5.7 Mass number	
8	14/11 - 21/11	Unit 5 Atomic structure 5.8 Isotopes 5.9 Relative masses of atoms 5.10 The arrangement of electrons in atoms Unit 6 Periodic Table Quiz on Names and Symbols of elements 6.1 How to group elements together? 6.2 The periodic table 6.3 Patterns across the periodic table	Ex.5
9	22/11 - 29/11	6.4 Group I elements 6.5 Group II elements 6.6 Group VII elements 6.7 Group 0 elements 6.8 Predicting the chemical properties of unfamiliar elements	Ex.6

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CYCLE	DATES	SYLLABUS TO BE COVERED	REMARKS
10	30/11 - 8/12	Unit 7 Ionic and metallic bonds 7.2 Chemical bonds 7.1 Conductors electrolytes and non-conductors 7.3 From atoms to ions	
11	12/12 - 18/12	Unit 7 Ionic and metallic bonds 7.4 Predicting the charge of an ion 7.5 Ionic bond (of sodium chloride, magnesium fluoride and lithium oxide)	Ex.7A
	3/1 - 16/1	First examination	1st exam
12	19/12 - 19/1	Discussion on Exam papers Unit 7 Ionic and metallic bonds 7.5 Compounds containing polyatomic ions 7.6 Names and formulae of ions	
13	22/1 - 29/1	Unit 7 Ionic and metallic bonds 7.6 Names and formulae of ions (continued) 7.7 Naming ionic compounds	
14	30/1 - 6/2	Unit 7 Ionic and metallic bonds Quiz on names and formulae of ions 7.9 Chemical formulae of ionic compounds	Ex.7B
15	7/2 - 28/2	Unit 7 Ionic and metallic bonds Quiz on names and formulae of ionic compounds 7.8 Colours of ionic compounds 7.9 Colours of gemstones 7.10 Movement of coloured ions 7.11 Metallic bonds	
16	29/2 - 12/3	Unit 8 Covalent bonds 8.1 Covalent bonds in non-metal elements	
17	13/3 - 20/3	Unit 8 Covalent bonds 8.2 Covalent compounds 8.3 Writing chemical formulae of covalent compounds	
18	8/4 - 15/4	Unit 8 Covalent bonds 8.4 Predicting the formation of ionic and covalent compounds	Ex. 8 2nd Uniform test
19	16/4 - 23/4	Unit 11 Reactivity of metals 11.1 Comparing the reactivity of metals 11.2 How do metals react with oxygen? 11.3 How do metals react with water or steam? 11.4 How do metals react with dilute acids?	
20	24/4 - 2/5	Unit 11 Reactivity of metals Experiment 11.1	
21	3/5 - 10/5	Unit 11 Reactivity of metals 11.6 What is chemical equation? 11.7 How to write balanced chemical equations	Ex.11A
22	13/5 - 21/5	Unit 11 Reactivity of metals Chemical equations for reactions in 11.2 to 11.4 11.8 What determines the reactivity of a metal? 11.9 Displacement reactions	

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23	22/5 - 29/5	Unit 11 Reactivity of metals Experiment 11.2 11.10 Ionic equations 11.11 Relationship between the extraction method and position of metals in the reactivity series 11.12 Reactivity series and reduction of metal oxides	Ex.11B
24	30/5 - 6/6	Revision	