PO LEUNG KUK CENTENARY LI SHIU CHUNG MEMORIAL COLLEGE
TEACHING SCHEDULE
FORM 3 CHEMISTRY
(2025-2026)

CYCLE	DATES	SYLLABUS TO BE COVERED	Values Education& National Security Education	REMARKS
1	2/9 – 10/9	Unit 1 Introducing Chemistry 1.1 What is Chemistry? 1.2 Why study Chemistry? 1.3 Laboratory safety 1.4 %Hazard warning labels 1.5 Common apparatus in the laboratory	%Law-abidingness	Ex.1
2	11/9 – 18/9	Unit 2 The atmosphere 2.1 The Earth 2.2 Classification of matter 2.3 Elements and compounds 2.4 Differences between a mixture and a compound 2.6 The atmosphere 2.7 Physical and Chemical Properties 2.8 Physical and chemical changes 2.9 Separation of mixtures 2.10 Separating oxygen and nitrogen from the air 2.11 Test for oxygen		Ex.2
3	19/9 – 26/9	Unit 3 The ocean 3.1 Sea water: a vast solution 3.2 What is a solution? 3.3 Obtaining common salt from sea water 3.4 %Obtaining pure water from sea water **Experiment 3.1**	% Responsibility	
4	29/9 – 8/10	Unit 3 The ocean 3.5 What does common salt contain? 3.6 Test for presence of water in a sample 3.7 Composition of sea water 3.8 Getting useful substances from sea water **Experiment 3.2**		Ex.3
5	9/10 - 16/10	Unit 4 : Rocks and minerals 4.1 Metals in the Earth's crust 4.2 Extracting metals from their ores 4.3 Investigating calcium carbonate		
6	17/10 - 3/11	Unit 4 The ocean 4.4 Formation of chalk, limestone and marbles 4.5 Formation of limestone caves **Experiment 4.1**		Ex.4 **1st Uniform test**
7	4/11 – 11/11	Unit 5 Atomic structure 5.1 What is an element made of? 5.2 Symbols for elements 5.3 States of elements 5.4 How to classify elements? 5.5 Basic Structure of an atom 5.6 Atomic number 5.7 Mass number		

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8	12/11 - 19/11	Unit 5 Atomic structure 5.8 Isotopes 5.9 Relative masses of atoms 5.10 The arrangement of electrons in atoms Unit 6 Periodic Table **Quiz on Names and Symbols of elements** 6.1 How to group elements together? 6.2 The periodic table 6.3 Patterns across the periodic table		Ex.5
9	24/11 -1/12	6.4 Group I elements 6.5 Group II elements 6.6 Group VII elements 6.7 Group 0 elements 6.8 Predicting the chemical properties of unfamiliar elements		Ex.6
10	2/12 - 16/12	Unit 7 Ionic and metallic bonds 7.2 Chemical bonds 7.1 Conductors electrolytes and nonconductors 7.3 From atoms to ions		
11	17/12 - 22/12	Unit 7 Ionic and metallic bonds 7.4 Predicting the charge of an ion 7.5 Ionic bond (of sodium chloride, magnesium fluoride and lithium oxide)		Ex.7A
	5/1 - 17/1	**First examination**		1st exam
13	19/1 - 20/1	Discussion on Exam papers Unit 7 Ionic and metallic bonds 7.5 Compounds containing polyatomic ions 7.6 Names and formulae of ions		
14	21/1 - 28/1	Unit 7 Ionic and metallic bonds 7.6 Names and formulae of ions (continued) 7.7 Naming ionic compounds		
15	29/1 - 5/2	Unit 7 Ionic and metallic bonds **Quiz on names and formulae of ions** 7.9 Chemical formulae of ionic compounds		Ex.7B
16	6/2 - 13/2	Unit 7 Ionic and metallic bonds **Quiz on names and formulae of ionic compounds** 7.8 Colours of ionic compounds 7.9 Colours of gemstones 7.10 Movement of coloured ions 7.11 Metallic bonds		
17	25/2 - 6/3	Unit 8 Covalent bonds 8.1 Covalent bonds in non-metal elements		
18	10/3 - 17/3	Unit 8 Covalent bonds 8.2 Covalent compounds 8.3 Writing chemical formulae of covalent compounds		
19	18/3 - 10/4	Unit 8 Covalent bonds 8.4 Predicting the formation of ionic and covalent compounds		Ex. 8 **2nd Uniform test**

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20	13/4 – 20/4	Unit 11 Reactivity of metals 11.1 Comparing the reactivity of metals 11.2 How do metals react with oxygen? 11.3 How do metals react with water or steam? 11.4 How do metals react with dilute acids?		
21	21/4 – 4/5	Unit 11 Reactivity of metals **Experiment 11.1**		
22	5/5 – 12/5	Unit 11 Reactivity of metals 11.6 What is chemical equation? 11.7 How to write balanced chemical equations		Ex.11A
23	13/5 – 20/5	Unit 11 Reactivity of metals Chemical equations for reactions in 11.2 to 11.4 11.8 What determines the reactivity of a metal? 11.9 Displacement reactions		
24	21/5 – 29/5	Unit 11 Reactivity of metals **Experiment 11.2** 11.10 Ionic equations 11.11 #Relationship between the extraction method and position of metals in the reactivity series 11.12 Reactivity series and reduction of metal oxides	% Responsibility #7 Understand the impact of human activities on the ecological environment and our responsibilities, understand the needs of sustainable development, and recognize the necessity of safeguarding ecological security, resource security, nuclear security and new security domains	Ex.11B
25	1/6 – 9/6	Revision		
	11/6 – 27/6	**Second examination**		2nd exam

Topic related to National Security Education
% Twelve priority values and attitudes:
Perseverance, Respect for Others, Responsibility, National Identity, Commitment, Integrity, Benevolence, Law-abidingness, Empathy, Diligence, Unity and Filial Piety